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Displayed equations should be centralised and numbered consecutively, with the equation number flush right (i.e. right-justified) and enclosed in parentheses. Equations should be numbered sequentially with Arabic numerals and referred to in the text as Eq. X. In multiple-line equations, the number should be given on the last

line. Please ensure that equations are numbered correctly, without repetition, and that no important equations are omitted from the numbering scheme. See Eq. 1 as an example of correct layout and numbering.

$$k_{et} = \kappa_{el} \nu_n \exp\left(-\frac{(\lambda + \Delta G^o)^2}{4\lambda RT}\right) \quad (1)$$

Equations should be typed in the same font size as the main text, with superscripts and subscripts 2–3 points smaller.

Conventional symbols and SI units should be adopted and used consistently.

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Tables should be placed in the text near where they are first referenced, with captions placed above the tables and sequentially numbered with Arabic numerals. Type the captions in 8 point, with a baselineskip of 11 points, see Table 1 as an example.

Table 1. Reduction and oxidation potentials for [ML*(NO)(X)(Y)] compounds, all redox potentials quoted vs SCE. L* = tris-(3,5-dimethylpyrazolyl)borato.

M	X	Y	E _{1/2} Reduction	E _{1/2} Oxidation
Mo	Cl	Cl	+0.09	—
W	Cl	Cl	-0.46	—
Mo	Cl	SPh	-0.37	—
W	Cl	SPh	-0.83	—
Mo	Cl	OPh	-0.42	—
W	Cl	OPh	-0.86	—
Mo	OPh	OPh	-0.74	—
W	OPh	OPh	-1.20	—
Mo	Cl	py	-1.60	+0.44
Mo	Cl	py-NMe ₂ -4	-2.25	0.00

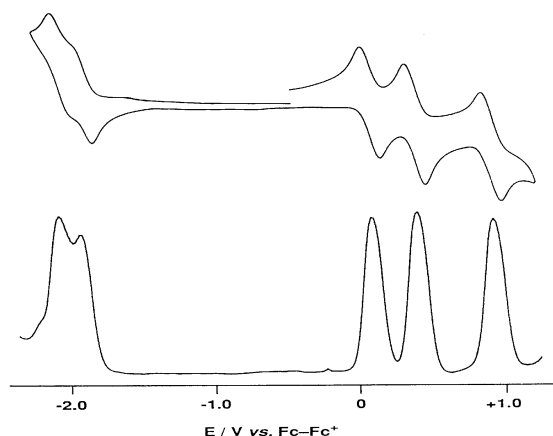


Figure 1. Cyclic and square-wave voltammograms of complex 33.

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1. Burke J. F. and Haare B. R., *Phys. Rev.* **D25** (1982), 2015–2017.
2. Harlee M. L. *et al.*, in *Superconductivity in d- and f-Band Metals*, ed. Davidsson D. H. (AIP, New York, 1972), 154–160.
3. Boot S., *Fundamentals of Inorganic Crystal Chemistry* (McGraw-Hill, London, 1968), 160–170.

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Acknowledgements

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Appendix

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